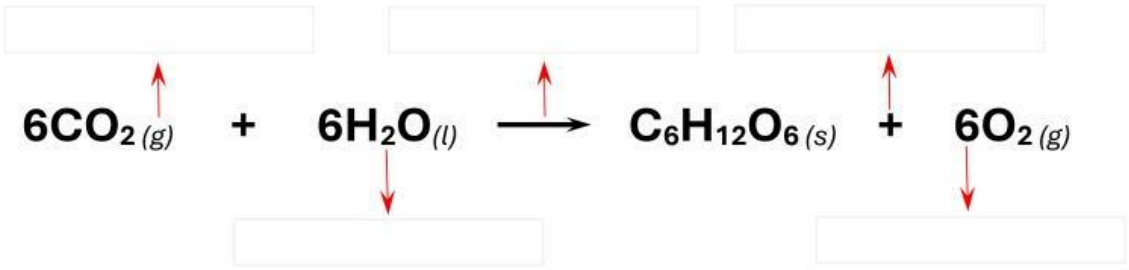
# Review CHAPTER 10

1. **Label the components of the Chemical Reaction below by selecting the correct options from the drop-down boxes :**




2. **Draw a line between the options on the left and the matching options on the right :**

| | |
|---|---|
| Products | Starting Materials |
| Precipitate | Number of Atoms |
| Subscript | Number of Molecules |
| Reactants | Newly formed Substances |
| Coefficient | Solid that forms in a liquid |

3. **Identify the type of Chemical Reaction in each of the following scenarios (select only 1 option in each case unless stated otherwise) :**

**3.1) Na + Cl -> Table salt**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

**3.2) $P_4$ + $O_2$ -> $P_2O_5$**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

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**3.3) $CH_4$ + $O_2$ -> $CO_2$ + $H_2O$ + Heat**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

**3.4) Zn + $CuSO_4$ -> $ZnSO_4$ + Cu (select 2 possible answers)**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Combustion
E) Redox

**3.5) Zn + $CuCl_2$ -> $ZnCl_2$ + Cu**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

**3.6) AX + BY -> AB + XY**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

**3.7) X + N + A + D -> ADNX**

A) Synthesis
B) Decomposition
C) Single Replacement
D) Double Replacement
E) Combustion

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4. **Drag and Drop the words from the Word Bank onto the correct places so that the statements below make sense :**

| gained | reactant | decrease | product | increase | lost |
|---|---|---|---|---|---|

In a Synthesis Reaction, 2 or more reactants are combined into a single _____ .

Oxygen is always a _____ in a Combustion Reaction.

In a Redox Reaction, the element that undergoes Reduction will have a(n) _____ in

Oxidation Number. This means that this element has _____ one or more Electrons.

The element that undergoes Oxidation, will have a(n) _____ in Oxidation Number.

This means that this element has _____ one or more Electrons.

5. **Balance the following Chemical Equations (The grading tool unfortunately cannot grade empty spaces, so please do NOT leave any spaces empty ... for the sake f the grading tool, please fill in the "1" in those spaces) :**

___ $Ba_3N_2$ + ___ HF -> ___ $BaF_2$ + ___ $NH_3$

___ $PCl_5$ + ___ $H_2O$ -> ___ $H_3PO_4$ + ___ HCl

___ $SnO_2$ + ___ $H_2$ -> ___ Sn + ___ $H_2O$

___ Al + ___ HCl -> ___ $AlCl_3$ + ___ $H_2$

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___ $KOH$ + ___ $H_3PO_4$ -> ___ $K_3PO_4$ + ___ $H_2O$

___ $HClO_4$ + ___ $P_4O_{10}$ -> ___ $H_3PO_4$ + ___ $Cl_2O_7$

___ $KNO_3$ + ___ $H_2CO_3$ -> ___ $K_2CO_3$ + ___ $HNO_3$

___ $CaS$ + ___ $H_2O$ -> ___ $Ca(HS)_2$ + ___ $Ca(OH)_2$

**6. Look at the Chemical Formulas provided and answer the questions that follow :**

**6.1) How many Carbon atoms are there in $C_4H_2S(OH)_2$ : _____**

**6.2) How many Carbon atoms are there in $3C_4H_2S(OH)_2$ : _____**

**6.3) How many Hydrogen atoms are there in $C_4H_2S(OH)_2$ : _____**

**6.4) How many Hydrogen atoms are there in $3C_4H_2S(OH)_2$ : _____**

**6.5) How many Oxygen atoms are there in $2C_4S(OH)_2$ : _____**

**6.6) How many Hydroxide atoms are there in $2C_4S(OH)_2$ : _____**

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