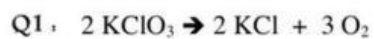


Cl = 35.45 a.m.u

K = 39 a.m.u

O = 16 a.m.u



How many moles of O_2 will be formed from 1.65 moles of KClO_3 ?

$$2.73 \text{ moles of } \text{KClO}_3 \times \frac{\text{mol } \text{O}_2}{\text{mol } \text{KClO}_3} = \text{mol } \text{O}_2$$

Q2 : How many **gram** of KCl will be formed from 2.73 moles of KClO_3 ?

$$2.73 \text{ moles of } \text{KClO}_3 \times \text{ } \times \text{ } = \text{gram of } \text{KCl}$$