



Section 4: Moles of Compounds -

Objective 1 - Convert between moles of a compound to mass.

What is the mass of 0.5 mol of H_2SO_4 ?

2- What is the mass, in grams, of potassium in 12.5 g of $\text{K}_2\text{Cr}_2\text{O}_7$ (molar mass = 294.40 g/mol)?

Challenge

What is the mass, in grams, of potassium in 12.5 g of $\text{K}_2\text{Cr}_2\text{O}_7$ (molar mass = 294.40 g/mol)?

2.50 Kg Fe_2O_3

Objective 2 : Convert between mass of a compound to mole

How many moles of copper are present in 148.2 g $\text{Ca}(\text{OH})_2$?

Determine the number of moles present in each compound:

22.6 g AgNO_3

35.0 g of hydrochloric acid (Challenge)