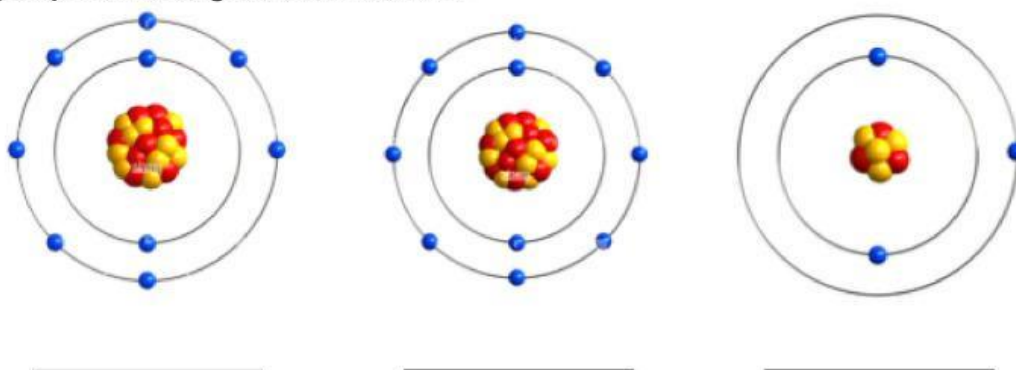


SCIENCE CLASS REVIEW / 9th grade

- By looking at valence electrons of the next Bohr's atomic models, select the group that belongs to each element.



- Indicate the kind of arrangement is showing each image of the periodic table.

Group number = number of outermost electrons

The image shows a periodic table with group numbers 1 through 8 indicated by red arrows. The groups are labeled as follows: 1 (H, Li, Na, K, Rb, Cs, Fr), 2 (He, Be, Mg, Ca, Sr, Ba, Ra), 3 (B, Al, Ga, In, Tl, Nh), 4 (C, Si, Ge, Sn, Pb, Fl), 5 (N, P, As, Sb, Bi, Mc), 6 (O, S, Se, Te, Po, Lv), 7 (F, Cl, Br, I, At, Ts), and 8 (He, Ne, Ar, Kr, Xe, Rn, Og). A red arrow points from the group number to the number of valence electrons in the outermost shell.

- Indicate next to each color if the elements belong to metalloids, metals, or nonmetals.

The image shows a periodic table with color-coded boxes for metalloids, metals, and nonmetals. The boxes are labeled as follows: 1 (H), 2 (He), 3 (Li), 4 (Be), 5 (B), 6 (C), 7 (N), 8 (O), 9 (F), 10 (Ne), 11 (Na), 12 (Mg), 13 (Al), 14 (Si), 15 (P), 16 (S), 17 (Cl), 18 (Ar), 19 (K), 20 (Ca), 21 (Sc), 22 (Ti), 23 (V), 24 (Cr), 25 (Mn), 26 (Fe), 27 (Co), 28 (Ni), 29 (Cu), 30 (Zn), 31 (Ga), 32 (Ge), 33 (As), 34 (Se), 35 (Br), 36 (Kr), 37 (Rb), 38 (Sr), 39 (Y), 40 (Zr), 41 (Nb), 42 (Mo), 43 (Tc), 44 (Ru), 45 (Rh), 46 (Pd), 47 (Ag), 48 (Cd), 49 (In), 50 (Sn), 51 (Sb), 52 (Te), 53 (I), 54 (Xe), 55 (Cs), 56 (Ba), 57 (La), 58 (Ce), 59 (Pr), 60 (Nd), 61 (Pm), 62 (Sm), 63 (Eu), 64 (Gd), 65 (Tb), 66 (Dy), 67 (Ho), 68 (Er), 69 (Tm), 70 (Yb), 71 (Lu), 72 (Hf), 73 (Ta), 74 (W), 75 (Re), 76 (Os), 77 (Ir), 78 (Pt), 79 (Au), 80 (Hg), 81 (Tl), 82 (Pb), 83 (Bi), 84 (Po), 85 (At), 86 (Rn), 87 (Fr), 88 (Ra), 89 (Ac), 90 (Th), 91 (Pa), 92 (U), 93 (Np), 94 (Pu), 95 (Am), 96 (Cm), 97 (Bk), 98 (Cf), 99 (Es), 100 (Fm), 101 (Md), 102 (No), 103 (Lr), 104 (Rf), 105 (Db), 106 (Sg), 107 (Bh), 108 (Hs), 109 (Mt), 110 (Ds), 111 (Nh), 112 (Fl), 113 (Mc), 114 (Lv), 115 (Ts), 116 (Og), 117 (Tennessine), 118 (Oganesson).

4. Select if the next characteristics belong to metals, nonmetals or metalloids.

• DULL • POOR CONDUCTORS • NONDUCTILE • BRITTLE • MAY BE SOLIDS, LIQUIDS OR GASES AT ROOM TEMPERATURE • TRANSPARENT AS A THIN SHEET • NOT SONOROUS	• LOOK LIKE METALS, BEHAVE LIKE NON-METALS. • ALL ARE SOLID AT ROOM TEMPERATURE • BRITTLE, HARD AND SOMEWHAT REACTIVE • INTERMEDIATE TO FAIRLY STRONG ELECTRICAL CONDUCTIVITY. • HAVE THE ABILITY TO FORM METALLIC ALLOYS.	• LUSTROUS • GOOD CONDUCTORS • HIGH DENSITY, HIGH MELTING POINT • MALLEABLE • DUCTILE (CAN BE DRAWN INTO WIRES) • USUALLY SOLID AT ROOM TEMPERATURE • OPAQUE AS A THIN SHEET • SONOROUS
--	--	---

5. Organize the next element if they are alkali metals, halogens or noble gases.

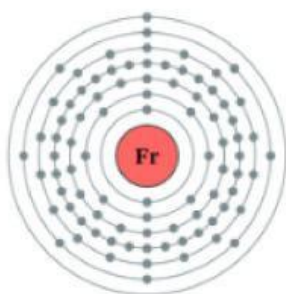
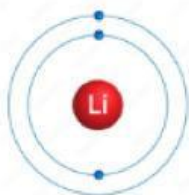
__Chlorine __Lithium __Fluorine
 __Xenon __Rubidium
__Sodium __Neon __Bromine
 __Cesium __Potassium
__Francium __Krypton
 __Iodine __Argon

6. According to the amount of reactivity, indicate which one is Potassium, Sodium, and Lithium.



7. The next groups are very reactive. Indicate per group which one is the most reactive and the less reactive.

Alkali metals



Halogens

