

1. A substance which oxidises itself and reduces other is known as -
- (a) oxidising agent (b) reducing agent
(c) both of these (d) none of these
2. Which of the following reactions involves the combination of two elements?
- (a) $\text{CaI} + \text{CO}_2 \rightarrow \text{CaCO}_3$ (b) $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
(c) $\text{SO}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{SO}_3$ (d) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
3. When hydrogen sulphide gas is passed through a blue solution of copper sulphate, a black precipitate of copper sulphide is obtained and the sulphuric acid so formed remains in the solution. The reaction is an example of-
- (a) a combination reaction
(b) a displacement reaction
(c) a decomposition reaction
(d) a double decomposition reaction
4. Which of the following is a physical change?
- (a) Formation of curd from milk
(b) Ripening of fruits
(c) Getting salt from sea water
(d) Burning of wood

5. What happens when copper rod is dipped in iron sulphate solution?
- (a) Copper displaces iron
 - (b) Blue colour of copper sulphate solution is obtained
 - (c) No reaction takes place
 - (d) Reaction is exothermic
6. A student added dilute HCl to a test tube containing zinc granules and made following observations :
- (a) the zinc surface became dull and black
 - (b) a gas evolved which burnt with a pop sound
 - (c) the solution remained colourless
 - (d) the solution becomes green in colour
7. A dilute solution of sodium carbonate was added to two test tubes - one containing dil HCl (a) and the other containing dilute NaOH(b). The correct observation was-
- (a) a brown coloured gas liberated in test tube A
 - (b) a brown coloured gas liberated in test tube B
 - (c) a colourless gas liberated in test tube A
 - (d) a colourless gas liberated in test tube B
8. A balanced chemical equation is in accordance with-
- (a) Avogadro's law
 - (b) law of multiple proportion
 - (c) law of conservation of mass
 - (d) law of gaseous volumes