

Wednesday 12/3

### **Type 1: Binary Ionic Compounds Formulas**

Contains two \_\_\_\_\_ —one \_\_\_\_\_ & one \_\_\_\_\_

1. Write the symbol and charge of the \_\_\_\_\_
2. Write the symbol and charge of the \_\_\_\_\_
3. Balance the \_\_\_\_\_ (to form a \_\_\_\_\_ compound) by using \_\_\_\_\_

### **Type 1: Binary Ionic Compounds Names**

Contains two elements—one \_\_\_\_\_ & one \_\_\_\_\_

1. Write the name of the \_\_\_\_\_ first, (the name doesn't change).
2. Write the name of the \_\_\_\_\_ second with “-\_\_\_\_\_” as the ending  
(subscripts do not matter in this type)

Thursday 12/4

### **Type #2: Polyatomic Ionic Compounds Formulas**

Contains at least one \_\_\_\_\_ ion

1. Write the symbol and charge of the \_\_\_\_\_ or polyatomic ion first
2. Write the symbol and charge of the \_\_\_\_\_ or polyatomic ion second
3. Balance the charges (to create a \_\_\_\_\_ compound) with \_\_\_\_\_
4. Use \_\_\_\_\_ when adding \_\_\_\_\_ to a polyatomic ion

### **Type #2: Polyatomic Ionic Compounds Names**

Contains at least one \_\_\_\_\_ ion

1. Write the name of the \_\_\_\_\_ or “\_\_\_\_\_” for NH<sub>4</sub>
2. Write the name of the polyatomic anion (do not change the ending) or the single element with “-ide”  
(subscripts within a polyatomic ion must match the name exactly. If there are parentheses, the polyatomic ion is \_\_\_\_\_ the parentheses)

**Friday 12/5**

### Multivalent Metal Compounds Formulas (Type #1 or #2)

The roman \_\_\_\_\_ indicates the \_\_\_\_\_ of the cation metal.

Follow the rules for Type #1 or Type #2 as it applies.

Type the **Roman Numeral** for each number below using I (eyes), and V's

|    |  |
|----|--|
| 1  |  |
| 2  |  |
| 3  |  |
| 4  |  |
| 5  |  |
| 6  |  |
| 7  |  |
| 8  |  |
| 9  |  |
| 10 |  |