

Balancing chemical equations Date: June ___th, 2025

Time: 40 minutes

Full Name: _____ Grade: 9___

Teacher's Name: Eliana Medrano A.

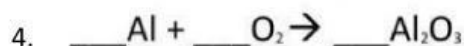
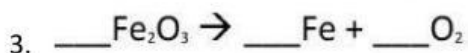
A. CIRCLE T (TRUE) F (FALSE)

	37.5
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1. A chemical formula of a compound provides the atomic number of the elements. T F
2. A chemical equation only shows the products. T F
3. The reactants are the result of a chemical reaction. T F
4. The reactants are separated by a plus sign. T F
5. The products of a chemical reaction are found to the right of the arrow. T F

B. IDENTIFY IF THE CHEMICAL EQUATIONS ARE BALANCED OR NOT.

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C. BALANCE THE CHEMICAL EQUATIONS (SHOW THE PROCESS)

37.5

Ex. 1. Na = $1 \times 2 = 2$

Cl = $1 \times 2 = 2$

O = $3 \times 2 = 6$

Na = $1 \times 2 = 2$

Cl = $1 \times 2 = 2$

O = $2 \times 3 = 6$

1. $\underline{\hspace{1cm}} \text{NaClO}_3 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{O}_2$
2. $\underline{\hspace{1cm}} (\text{NH}_4)_3\text{PO}_4 + \underline{\hspace{1cm}} \text{Pb}(\text{NO}_3)_4 \rightarrow \underline{\hspace{1cm}} \text{Pb}_3(\text{PO}_4)_4 + \underline{\hspace{1cm}} \text{NH}_4\text{NO}_3$
3. $\underline{\hspace{1cm}} \text{BF}_3 + \underline{\hspace{1cm}} \text{Li}_2\text{SO}_3 \rightarrow \underline{\hspace{1cm}} \text{B}_2(\text{SO}_3)_3 + \underline{\hspace{1cm}} \text{LiF}$
4. $\underline{\hspace{1cm}} \text{C}_7\text{H}_{17} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
5. $\underline{\hspace{1cm}} \text{CaCO}_3 + \underline{\hspace{1cm}} \text{H}_3\text{PO}_4 \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{H}_2\text{CO}_3$
6. $\underline{\hspace{1cm}} \text{Ag}_2\text{S} \rightarrow \underline{\hspace{1cm}} \text{Ag} + \underline{\hspace{1cm}} \text{S}_8$