Particle, Molar Mass, Mole and % Composition Quiz Retake

Name_____

USE YOUR UNITS AND SHOW YOUR WORK

1. A bottle contains 3.88 moles of carbon dioxide. How many molecules of carbon dioxide are in the bottle?

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2. A tube contains 5.44×10^{21} atoms of neon. How many moles of neon are in the tube?

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3. Fill in the table for the following substances?

Name	Formula	Molar Mass
bromine gas		
magnesium carbonate		
	$Ga_2(Cr_2O_7)_3$	
	PI_3	

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4. What is the percent composition of nickel(IV) nitride?

Formula_____

Molar Mass_____

Add as many tables as you need for each element

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5. What is the percent by mass of **hydroxide** in cobalt(III) hydroxide?

Formula_____

Molar Mass_____

=

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6. A beaker contains 22.3 g of zinc bromate. How many mol of zinc bromate are in the beaker?

Formula_____

=

7. A student has a flask that contains 0.58 mol of magnesium sulfate. How many grams of magnesium sulfate are in the flask?

Formula_____

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