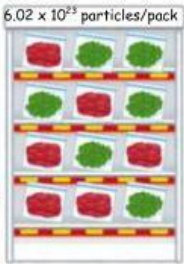
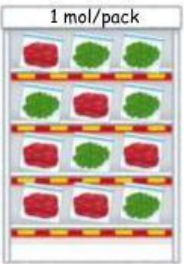
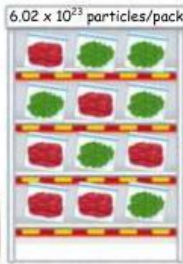


UNDERSTANDING THE MOLE AND AVOGADRO'S CONSTANT

A chemist and a physicist went shopping for Carbon-12. Help them complete their tasks as they need!

$$1 \text{ mol} = 6.02 \times 10^{23} \text{ particles} = 12 \text{ g of Carbon-12}$$

The store sells carbon-12 in different quantities according to the label on the shelf. Type in the needed amount on their baskets.

I need 2 mol of carbon  	I need 30.1×10^{23} particles  	I need 12 g of carbon  
I need 36 g of carbon  	I need 1 mol of carbon  	I need 6.02×10^{23} particles  