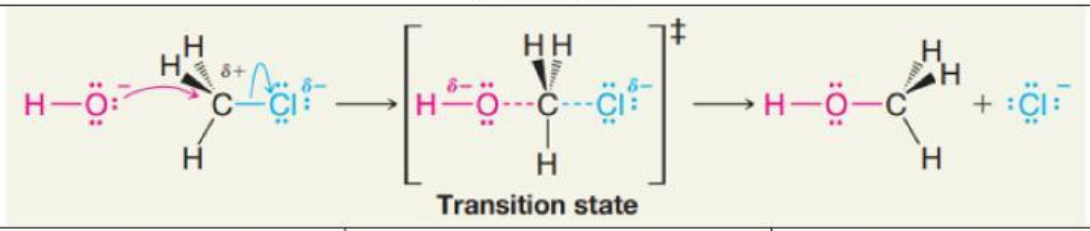
Fill in the blanks:

Mechanism for the S_N2 Reaction

	$HO^- + CH_3Cl \longrightarrow CH_3OH + Cl^-$		
Step	Mechanism		
			
	The negative hydroxide ion brings a pair of electrons to the partically positive *carbon/hydrogen* from the back side with respect to the leaving group. The chlorine begins to move away with the pair of electrons that bonded it to the *carbon/hydrogen*.	In the transition state, a bond between oxygen and carbon is partially *broken/formed* and the bond between carbon and chlorine is partially *broken/formed*. The configuration of the carbon atom begins to invert.	Now the bond between the oxygen and carbon has *departed/formed* and the chloride ion has *departed/formed*. The configuration of the carbon has inverted.