

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Chemical Bonding Worksheet

**REMEMBER....**

<b>Ionic Bond</b>	between a Metal and Non-Metal	<b>(M + NM)</b>
<b>Covalent Bond</b>	between a Non-Metal and Non-Metal	<b>(NM + NM)</b>

**PART 1:** Determine if the elements in the following compounds are metals or non-metals. Describe the type of bonding that occurs in the compound.

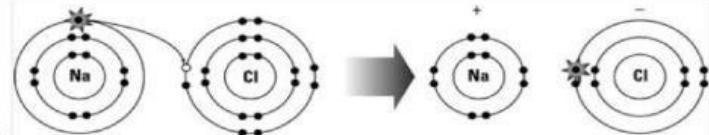
Compound	Element 1	M or NM	Element 2	M or NM	Bond Type
NO <sub>2</sub>	Nitrogen	NM	Oxygen	NM	Covalent
NaCl					
SO <sub>2</sub>					
PI <sub>3</sub>					
MgBr <sub>2</sub>					
CaO					
H <sub>2</sub> O					
K <sub>2</sub> O					
AlF <sub>3</sub>					
O <sub>2</sub>					
CuCl <sub>2</sub>					
NO <sub>2</sub>					
CO <sub>2</sub>					
HF					
Rb <sub>2</sub> S					
NBr <sub>2</sub>					
Fe <sub>2</sub> O <sub>3</sub>					
CCl <sub>4</sub>					

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## Chemical Bonding Worksheet

**PART 2:** Determine whether the following are properties of Ionic or Covalent Bonds.

<b>Chemical Property</b>	<b>Type of Bond</b>
1. Attraction between oppositely charged ions	Ionic
2. High Melting Point	
3. Low Melting Point	
4. High Boiling Point	
5. Low Boiling Point	
6. Transfer electrons	
7. Share Electrons	
8. Form crystal lattice	
9. Form Molecules	
10. Does not conduct electricity	
11. Conducts electricity in molten state (liquid state)	
12. Typically solid at room temperature	
13. Hard, brittle solid with definite crystal shape	
14. Dissociate into ions when dissolved in water	
15. Prefixes used when naming compounds	
16. Roman Numeral may be used when naming compounds	
17. Metal + Non-metal	
18. Non-metal + Non-metal	
19. Uses 'crisscross' method when writing formulas	
20. Uses prefixes to determine number of atoms in formula	
21. Typically liquid, gas, or soft solid at room temperature	
22. Solubility in water depends on polarity (polar or non-polar)	
23. Cations and Anions	
24. 	
25. 