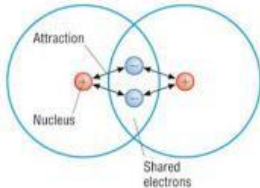
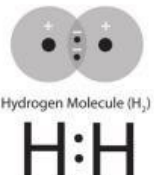
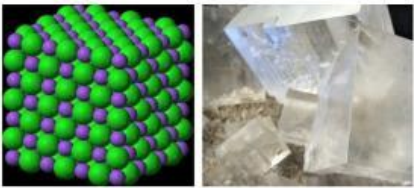
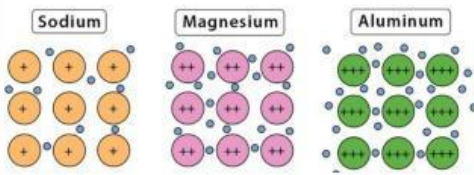

Cations and anions are formed (transfer of electrons).	Electrons are shared between atoms / tend to be between atoms.	Electrons are released from atoms and can move freely.
nonmetal + nonmetal	metal + metal	metal + nonmetal
Opposite charges attract (electrostatic forces between ions)	Two nuclei pull on shared electrons	“sea of electrons” or “electron glue”
Can be bent without breaking (think bending a copper wire)	If substance is solid, it may break (think breaking ice or a plastic cup)	Breaks easily (think smashing a salt crystal)
Individual molecules are formed	Conducts electricity	Lattice (tight packing of ions)
<b>Examples:</b> NaCl, CaBr <sub>2</sub>	<b>Examples:</b> Steel wire, aluminum foil	<b>Examples:</b> H <sub>2</sub> O, N <sub>2</sub> (nitrogen gas), DNA
 		

**ATR (\*) Types of Bonds** - Cut out the tiles / pictures on the next page and sort them into the table below.

Ionic Bond	Covalent Bond	Metallic Bond
1st picture or diagram here	1st picture or diagram here	1st picture or diagram here
<b>Example Tile here</b>	<b>Example tile here</b>	<b>Example tile here</b>
2nd picture or diagram here	2nd picture or diagram here	2nd picture or diagram here