

Chemical Equation

Name :

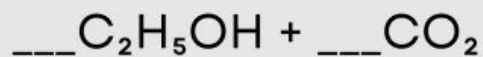
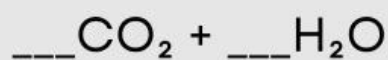
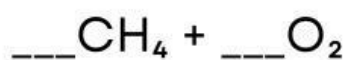
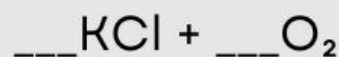
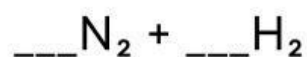
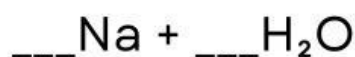
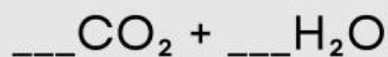
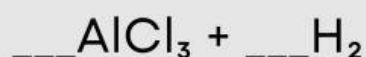
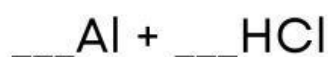
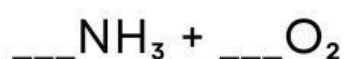
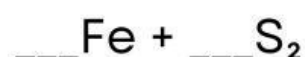
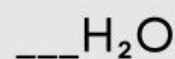
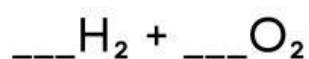
Class :

Balance the chemical equation below!

1	$\dots \text{H}_2 + \dots \text{Br}_2 \longrightarrow \dots \text{HBr}$
2	$\dots \text{N}_2 + \dots \text{H}_2 \longrightarrow \dots \text{NH}_3$
3	$\dots \text{K} + \dots \text{H}_2\text{O} \longrightarrow \dots \text{KOH}$
4	$\dots \text{S}_8 + \dots \text{O}_2 \longrightarrow \dots \text{SO}_2$
5	$\dots \text{H}_2\text{O}_2 \longrightarrow \dots \text{H}_2\text{O} + \dots \text{O}_2$
6	$\dots \text{P}_4\text{O}_{10} + \dots \text{H}_2\text{O} \longrightarrow \dots \text{H}_3\text{PO}_4$
7	$\dots \text{Zn} + \dots \text{AgCl} \longrightarrow \dots \text{ZnCl}_2 + \dots \text{Ag}$
8	$\dots \text{Al} + \dots \text{H}_2\text{SO}_4 \longrightarrow \dots \text{Al}_2(\text{SO}_4)_3 + \dots \text{H}_2$
9	$\dots \text{NaHCO}_3 \longrightarrow \dots \text{Na}_2\text{CO}_3 + \dots \text{H}_2\text{O} + \dots \text{CO}_2$
10	$\dots \text{CaCO}_3 + \dots \text{HCl} \longrightarrow \dots \text{CaCl}_2 + \dots \text{H}_2\text{O} + \dots \text{CO}_2$

Balancing Chemical Equations

Balance the following chemical equations by adjusting the coefficients in front of each compound. Make sure the number of atoms of each element on both sides of the equation is the same.

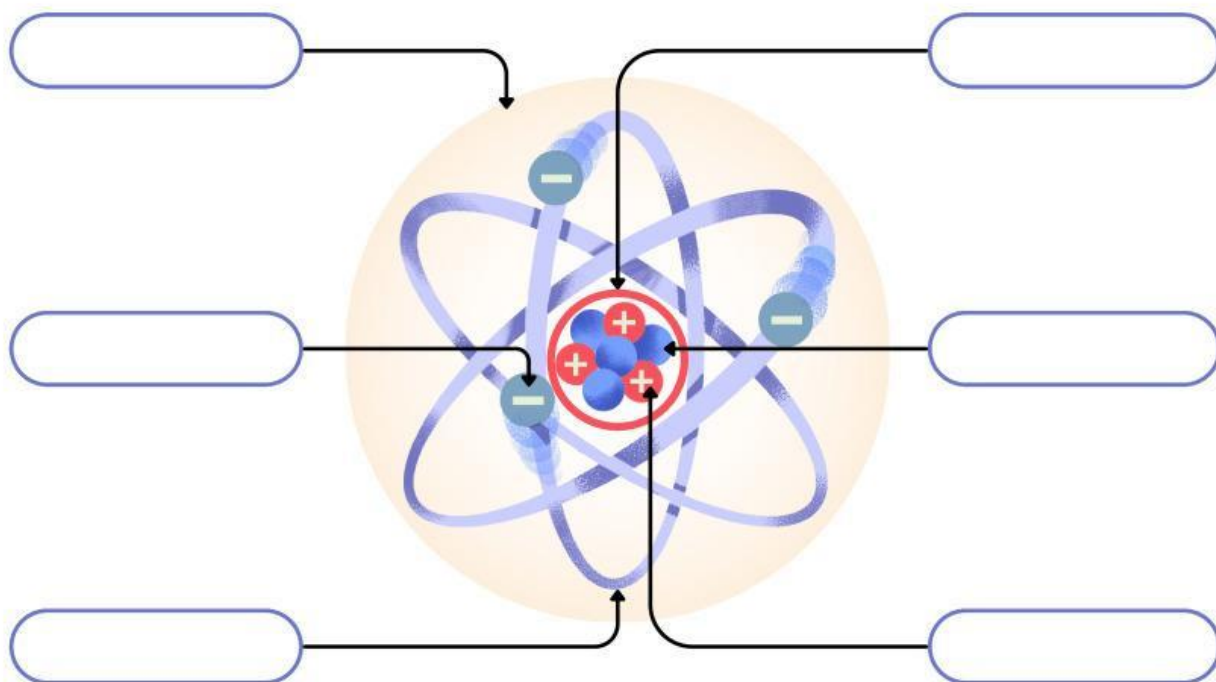


ATOMIC STRUCTURE

Name:

Date:

Exercise 1: Label the structures indicated with their respective names.



Exercise 2: Match the names to the structures described in the following sentences.

Part of the atom that has no charge	
Part of the atom that has a positive charge	
The atomic number is determined by this	
It must have the same number of these particles	
Part of the atom that has a negative charge	
The mass number is determined by this	