

STRONG BASE

Given the concentration of $\text{NaOH} = 0.2 \text{ M}$.
Calculate the pH of of the base solution

$$\begin{aligned} \text{pOH} &= -\log [\text{OH}^-] \\ &= -\log \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}} \end{aligned}$$

$$\begin{aligned} \text{pH} &= 14 - \text{pOH} \\ &= 14 - \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}} \end{aligned}$$

STRONG BASE

Given the concentration of $\text{Ba}(\text{OH})_2 = 0.2 \text{ M}$.

Calculate the pH of of the base solution

$$\text{pOH} = -\log [\text{OH}^-]$$

$$= -\log \underline{\hspace{2cm}}$$

$$= \underline{\hspace{2cm}}$$

$$\text{pH} = 14 - \text{pOH}$$

$$= 14 - \underline{\hspace{2cm}}$$

$$= \underline{\hspace{2cm}}$$