

### Q3 AT Mole worksheet

Name \_\_\_\_\_

Date \_\_\_\_ / \_\_\_\_ / \_\_\_\_

Pd \_\_\_\_

**Directions:** For all problems you must show the formula of the compound, your work with units and your final answer with units to receive full credit.

1. How many moles are represented by 54.87g of nitrogen pentachloride?

Formula: \_\_\_\_\_

Answer \_\_\_\_\_

2. How many grams would you need to weigh out to have 3.75 moles of gold(II) nitrate?

Formula: \_\_\_\_\_

Answer \_\_\_\_\_

3. How much would 0.67 mole of iron(V) oxide weigh in grams?

Formula: \_\_\_\_\_

Answer \_\_\_\_\_

4. How many moles of sodium sulfite are in 245.30g?

Formula: \_\_\_\_\_

Answer \_\_\_\_\_

5. How many moles are in 874.21g of silicon octoxide?

Formula: \_\_\_\_\_

Answer \_\_\_\_\_