

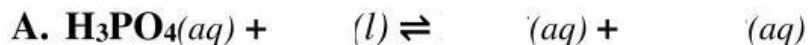
Elite Private School

Worksheet

ACID AND BASE DISSOCIATION.

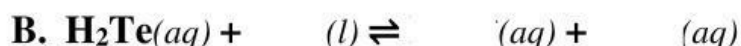


Using the correct information from each table, complete the chemical equation and dissociation expressions for each **ACID**.



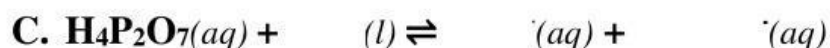
$K_a = \underline{\hspace{2cm}}$

H_3O^+	H_2O	H_2PO_4^-
$[\text{H}_3\text{O}^+]$	$[\text{H}_3\text{PO}_4]$	$[\text{H}_2\text{PO}_4^-]$



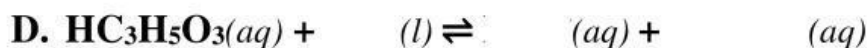
$K_a = \underline{\hspace{2cm}}$

H_3O^+	H_2O	HTe^-
$[\text{H}_3\text{O}^+]$	$[\text{H}_2\text{Te}]$	$[\text{HTe}^-]$



$K_a = \underline{\hspace{2cm}}$

H_3O^+	H_2O	$\text{H}_3\text{P}_2\text{O}_7^-$
$[\text{H}_3\text{O}^+]$	$[\text{H}_4\text{P}_2\text{O}_7]$	$[\text{H}_3\text{P}_2\text{O}_7^-]$



$K_a = \underline{\hspace{2cm}}$

H_3O^+	H_2O	$\text{C}_3\text{H}_5\text{O}_3^-$
$[\text{H}_3\text{O}^+]$	$[\text{HC}_3\text{H}_5\text{O}_3]$	$[\text{C}_3\text{H}_5\text{O}_3^-]$

E. Write the chemical equation and dissociation expression for the weak acid, H_2SO_3 .

F. Write the following:

- a. The equation for the acid H_2CO_3 ;**
- b. The equation for the conjugate base of the acid**
- c. The K_a expression for the acid**
- d. The K_b expression for the conjugate base.**