

Name \_\_\_\_\_ Period \_\_\_\_\_

**Directions:** Write and balance the chemical equation (including states) indicated by the sentence and indicate the type of chemical reaction.

**1.) Aqueous silver nitrate is mixed with aqueous calcium chloride to form solid silver chloride and aqueous calcium nitrate.**

Equation:

Type of reaction:

**2.) Solid sodium reacts with aqueous iron(III) chloride to produce solid iron and aqueous sodium chloride.**

Equation:

Type of reaction:

**3.) Water reacts with carbon dioxide gas to produce aqueous carbonic acid( $\text{H}_2\text{CO}_3$ ).**

Equation:

Type of reaction:

**4.) Gaseous pentane( $\text{C}_5\text{H}_{12}$ ) reacts with oxygen gas to produce carbon dioxide gas and water vapor.**

Equation:

Type of reaction:

**5.) Aqueous zinc chloride reacts with dihydrogen monosulfide gas to yield a zinc sulfide precipitate and hydrochloric acid( $\text{HCl}(\text{aq})$ )**

Equation:

Type of reaction:

**6.) Solid iron(III)oxide reacts with solid aluminum metal to produce solid aluminum oxide and solid iron.**

Equation:

Type of reaction:

**7.) Carbon monoxide gas reacts with oxygen gas to produce carbon dioxide gas.**

Equation:

Type of reaction:

**8.) Solid sodium hydroxide breaks down into solid sodium oxide and water vapor.**

Equation:

Type of reaction:

**9.) Aqueous nickel(III) nitrate reacts with chromium metal to produce chromium(II) nitrate solution and solid nickel.**

Equation:

Type of reaction:

**10.) Solid iron(III) oxide reacts with water to form aqueous iron(III) hydroxide.**

Equation:

Type of reaction: