

NAME : _____

CLASS : _____

DATE : _____

1. Alkali metals are found in which group of the periodic table?

Group 1

Group 2

Group 7

group 8

2. Which of the following is the most reactive in water?

Lithium

Sodium

Potassium

3.

Alkali are so reactive to air they are generally stored in what?

Water

Oil

Vaults

Cotton

a vacuum

4. ----- is released when alkali metal react with water

Hydrogen

Oxygen

Metahne

Carbon di oxide

5. What happens to the reactivity of the alkali metals as you move down the group?

increases

decreases

6. Which alkali metal is the least reactive?

lithium

sodium

potassium

rubidium

7. Which of the following is not an alkali metal

Sodium

Caesium

Francium

Calcium

Potassium

8. Alkali earth metals such as Magnesium produce when they are burned.

very bright light

huge explosion

hydrogen gas

salt

9. Alkali earth metals are found in which group of the periodic table?

Group 1

Group 2

Group 7

group 8

10. Which of the following is NOT an alkali earth metal?

Be

Mg

Ca

Sn

11. An element in Group 2 of the periodic table.

Alkali metal

Alkaline Earth metal

Transition metal

Metalloid

12.



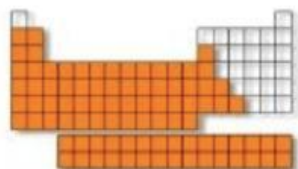
What category is highlighted?

metals

nonmetals

metalloids

13.



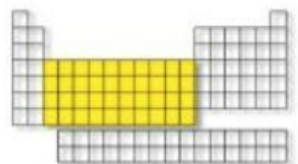
What category is highlighted?

metals

nonmetals

metalloids

14.



Which group is highlighted?

Alkali Metals

Halogens

Transition Metals

15.



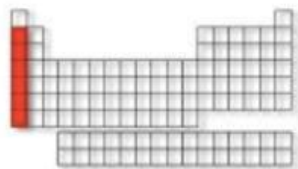
What category is highlighted?

metals

nonmetals

metalloids

16.



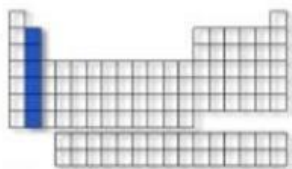
What group is highlighted?

Transition Metals

Alkali Metals

Carbon Family

17. What group is highlighted?



Alkaline Earth Metals

Transition Metals

Nitrogen Family

18. What group is highlighted?



Halogens

Noble Gases

Lanthanides

19. Electrons on the outer most energy level of an atom.

valence electron

noble gasses

electrons

electron cloud

20. Which gas is released when the alkali metals react with water?

hydrogen

oxygen

hydroxide

carbon dioxide

21. How does a group 1 metal atom form an ion?

lose 1 electron

lose 2 electrons

gain 1 electron

gain 2 electrons