Dalton's Law of Partial Pressure

(make sure you round to the nearest hundreths decimail place)

1. CO_2 , O_2, and H_2 are placed in a container. The gases are at standard temperature and pressure. The pressure for CO_2 is 2.5 atm, the pressure for O_2 is 0.7 atm, and the pressure of H_2 is 1.2 atm. What is the total pressure of the mixture?

2. Two unknown gases are placed in a in a container at standard temperature and pressure. Unknown gas A has a pressure of 800 mmHg. The total pressure in the container for both gases is 1125 mmHg. What is the partial pressure for unknown gas B?

3. 3.2 atm of N_2 and 0.5 atm of CO are placed in a container at STP. What is the total pressure?

4. 900 mmHg of Cl_2 gas is mixed with NO_2 gas. The total pressure of the mixture is 2225mmHg. What is the partial pressure of NO_2 gas in the mixture?

5. 1.3 atm of gas A, 522 mmHg of gas B, and 78.3 kPa of gas C are placed in a container at standard temperature and pressure. What is the total pressure of the mixture?(answer should be in kPa)

6. 88.7 mmHg of N_2, 3.7 atm of CO, 900 torr of F_2, and Cl_2 gas are placed in a container. The total pressure for the mixture is 7.2 atm. What is the partial pressure of Cl_2 gas in the mixture? (answer should be in atm)

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