Name:_____Date:_____Period:_____

Task: Use the Law of Conservation of Mass to fill out the missing information in the table below. Use example #1 as a guide.

Mass of Reactants = Mass of Products

REACTION	Reactant(s)		Product(s)	
1	H_2	O_2	H_20	
mass	3.4g	10g	13.4g	
2	CH	O_2	CO_2	H_2O
mass	14.4g	40.0g		15.0g
3	HgO		Hg	O_2
mass	33.6g			18.3g
4	Li	O_2	Li_2O	
mass		10.8g	34.6g	
5	C_3H_6	O_2	CO_2	H_2O
mass	20.0g		34.2g	12.8g
6	$Al(OH)_3$		Al_2O_3	H_2O
mass			41.8g	15.6g
7	Zn	HCl	$ZnCl_2$	H_2
mass	34g	55g		10g
8	H_2O_2		H_2O	O_2
mass	85g			44g
9	CH_4	O_2	CO_2	H_2O
mass	120g	33g		65g
10	Na	Cl_2	NaCl	
mass	44g	45g		

Answer the word problems below using the Law of Conservation of Mass.

11. Hydrogen and oxygen react chemically to form water. How much water would form if 24.6 grams of hydrogen reacted with 33.6 grams of oxygen? ($H_2 + O_2 \rightarrow H_2O$)

12. When ammonium nitrate (NH_4NO_3) explodes, the products are nitrogen, oxygen, and water. When 55 grams of ammonium nitrate explode, 22 grams of nitrogen and 10 grams of oxygen are formed. How many grams of water is produced? ($NH_4NO_3 \rightarrow N_2 + O_2 + H_2O$)

LIVEWORKSHEETS