

Ionic Bonding

Binary Ionic Nomenclature

Ionic bonds are bonds that involve a **METAL** and a **NONMETAL**.

When writing the chemical formulas for ionic bonds we must:

1. Find the oxidation number/charge for each element
2. The metal/cation always comes first
3. Cross the oxidation numbers/charges so that they are now subscripts on the other element.
 - a. Do not record 1 as a subscript.
 - b. Reduce subscripts if possible.

*For the chart below write the **chemical formula** for each ionic bond.

	C	N	P	S	F
Li	Li ₄ C				
Ca					
K					
Sr					
Al					

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When naming ionic bonds we must first write the name of the metal and then the name of the nonmetal. The ending of the nonmetal needs to be changed to -ide.

Example: K and Se are bonded together. We would call this **potassium selenide**.

* In the chart below write the names of each ionic bond.

	C	N	S	Cl	F
Li	lithium carbide				
Ca					
K					
Sr					
Al					