

18. Saline solution is prepared by dissolving 9.0 g of NaCl in deionized water in a 500 ml volumetric flask. Calculate the molarity of the solution.

A. 3.07 M C. 0.307 M
B. 30.07 M D. 7.03 M

19. Given equations below:



How many mole of $\text{S}_2\text{O}_8^{2-}$ are needed to oxidise 20mL, 0.2 M Mn^{2+} ?

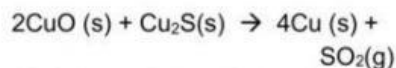
A. 4×10^{-3} mol C. 1×10^{-2} mol
B. 1×10^{-3} mol D. 2×10^{-2} mol

20. A compound has an molecular formula of $\text{C}_4\text{H}_5\text{ON}_2$. Chose the **correct** statement.

- I. The weight percentage of Hydrogen is 10.15%
II. The number of Nitrogen atoms is 1.204×10^{24}
III. The weight percentage of Carbon is 49.48%

A. I only C. I and III
B. I and II D. II and III

24. Based on balance equation below, determine the number of mole of excess reactant remained at the end of the reaction if 90.0 g of CuO is heated with 150.0 g Cu_2S .



A. 0.375 mol C. 0.565 mol
B. 1.13 mol D. 0.190 mol

25. The density of 10.5 molal NaOH is 1.33 g/mL. Calculate percentage by mass of NaOH.

A. 12.67% C. 29.57%
B. 75.18% D. 8.64%

27. 10 cm^3 of ammonia solution, NH_3 with density of 0.93 g/cm^3 contains 0.45 g of NH_3 solute. What is the percentage by mass of this solution?

A. 4.18 C. 2.22
B. 4.84 D. 4.20