

ATOMIC AND IONIC ISOTOPIC NOTATION

Fill in the blank table and figure out the last two isotopic notation.

Isotopic Notation	Number of proton	Number of neutrons	Number of electrons	Nucleon number
${}^1_1\text{H}$				
${}^1_1\text{H} +$				
${}^{35}_{17}\text{Cl} -$				
${}^{24}_{12}\text{Mg} 2 +$				
${}^{108}_{47}\text{Ag} +$				
${}^{32}_{16}\text{S} 2 -$				
	30		28	66
	76	114		