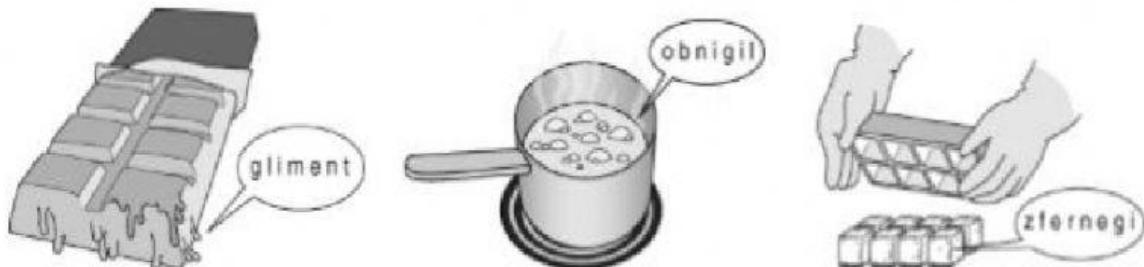
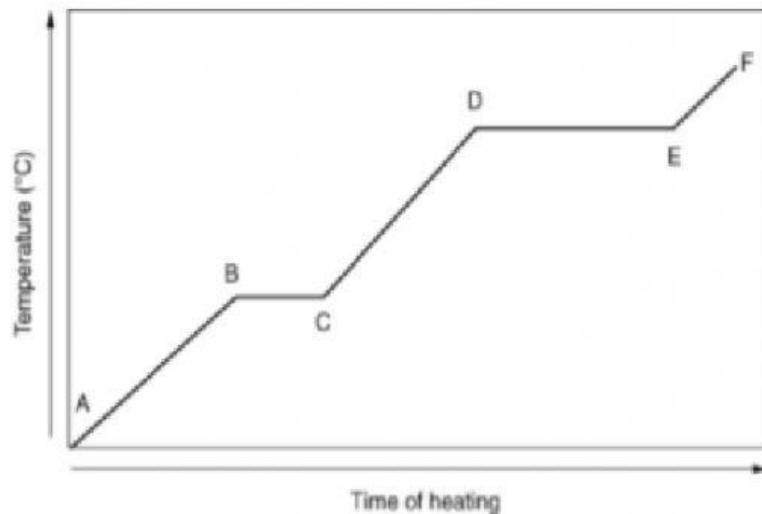


Arrange the jumbled letters to make a word that describes the change shown in the drawings.



Look at the graph shown below, and then answer the questions which follow.



- a What is happening to the temperature in section E to F?
- b What is the state of the substance in section A to B?
- c What change of state is occurring in section D to E?
- d Explain why the temperature doesn't rise in section B to C although heating continues.

The melting points of elements in the periodic table are shown below.

H -259.1						
Li 180.5						He -268.93
Na 97.6						
K 63.5	Ca 842					
Rb 39.3	Sr 777					
Cs 28.44	Ba 727					
Fr 27	Ra 606					
		B 2075	C 3625 sp	N -210	O -218.79	F -219.00
						Ne -248.609
		Al 660.32	Si 1414	P 44.15	S 115.21	Cl -101.5
						Ar -189.36
		Ga 29.76	Ge 617	As 817	Se 221	Br -157.36
						Kr -111.34
		In 156.6	Sn 231.93	Sb 630.63	Te 449.51	I 153.7
						Xe -111.34
		Tl 364	Pb 327.46	Bi 271.4	Po 554	At 907
						Rn -71

- a Name four elements in the halogen group.
- b Describe the general trend in melting points down the alkali metal and noble gas groups.
- c What is the general trend in melting points down group 6 from oxygen to polonium? Identify any melting point within this group that does not seem to fit the trend.
- d The melting points of germanium and bromine are missing. Estimate these melting points and explain your reasoning.

Tick (✓) the correct boxes to show the typical properties of metals and non-metals.

	Good conductors of electricity	Most are brittle when solid	Most have high melting points
Metals			
Non-metals			