Le Chatelier's Equilibrium Practice

Fill out the chart using the following equation based on the rules of Le Chatelier's Principle:
*The ΔH is _____ so this is an _____ reaction, so I will write the word HEAT on the _____ with the _____.

*There are more moles of gas on the _____ side of the arrow so I will write the word PRESSURE on the _____ with the _____ and I will write the word VOLUME on the opposite side.

$C_3H_{8(g)} + 5O_{2(g)} <--> 3CO_{2(g)} + 4H_2O_{(g)}$ $\Delta H = -45kJ$

Stress	Equilibrium Shift L or R	$[C_3H_8]$ increase or decrease	$[O_2]$ increase or decrease	$[CO_2]$ increase or decrease	$[H_2O]$ increase or decrease	Temp/Heat increase or decrease
Add O_2						
30°C → 15°C						
Remove C_3H_8						
Increase pressure						
Increase carbon dioxide						
4L → 8L						
Remove water						
Heat added						
Decrease pressure						