

Learning Target: I can use the International Union of Pure and Applied Chemistry (IUPAC) nomenclature for translating between simple binary chemical names and chemical formulas (one to one chemical compounds). (Proficient)



Writing Binary Ionic Compounds & Formulas 101 Video Review

1. What is an ion? _____
2. An anion has a _____ charge, while a cation has a _____ charge.
3. What is a Binary ionic compound? _____

4. What do you need to know in order to write binary ionic compounds? _____

5. We need to determine the _____ and make sure _____ cancel out to ____.
6. Write the example for Na & Cl below: _____ Write the example for Ca & O below: _____
7. What is the cross method? _____
8. If Ca has a +2 charge and Cl has a -1 charge what does this mean? _____

9. Write the cross method for Ca & Cl below.

+1		+2				+3		+4		-3		-2		-1		0
IA		IIA		IIIA	IVA	VA	VI	VIIA	VIIIA							
H		Li	Be	B	C	N	O	F	Ne							
Na		Mg		Al	Si	P	S	Cl	Ar							
K		Ca	Sc	Ti												
Rb		Sr	Y	Zr												
Cs		Ba	La	Hf												
Fr		Ra	Ac	Rf												

10. Write the cross method for each of the following chemical formulas.

Calcium & Sulfur = _____ - Lithium & Oxygen = _____ - Aluminum & Fluorine = _____

11. What 2 things do you need to do to name binary ionic compounds?

a. _____

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b. _____

Example: NaCl = _____

12. CaF_2 = _____ NaCl = _____ CaCl_2 _____

CaO = _____

13. Transition metals can have _____

14. Write the example for Iron (Fe) below:

15. Practice - Iron (III) = +3 Oxygen = -2 Chemical formula = _____ Chemical name = _____

16. Titanium (III) Nitride = _____ Iron (III) chloride = _____ Iron (III) sulfide = _____

Quiz – Write the chemical name for 1-5, and the chemical formula for 6-10. You have 2 minutes.

1. MgF_2 _____

2. ZnCl_2 _____

3. K_2O _____

4. Al_2O_3 _____

5. Na_2S _____

6. Aluminum Sulfide _____

7. Iron (III) Bromide _____

8. Lithium Oxide _____

9. Beryllium Chloride _____

10. Titanium (III) Phosphide _____