Calculate the Average Atomic Mass for the following FAKE elements:
Follow this format: Use parenthesis to multiply numbers, include units, do NOT leave spaces, round to the correct number of significant figures.

1. The element **BLUE** has naturally occurring isotopes with mass numbers of 63 and 65.
The relative abundance and atomic masses are:

69.2% for a mass of 85.93 amu
30.8% for a mass of 81.03 amu

Calculate the average atomic mass of BLUE.
Show work:

Final Answer:

2. **Calculate the average atomic mass of RED** if 95.00% of all the atoms have a mass of 11.972 amu, 0.76% has a mass of 12.971 amu and 4.22% have a mass of 13.967 amu.
Show work:

Final Answer:

3. Naturally occurring **GREEN** consists of four isotopes, Gr-54, Gr-56, Gr-57 and Gr-58.
Below is the data concerning strontium: What is the average atomic mass for **GREEN**?

Gr-54	Gr-56	Gr-57	Gr-58
0.56%	9.86%	7.00%	82.58%

Show work:

Final Answer:

4. **Calculate the average atomic mass of PURPLE**. One isotope of **PURPLE** has an atomic mass of 98.92amu and a relative abundance of 50.69%. The other major isotope of **PURPLE** has an atomic mass of 90.92amu and a relative abundance of 49.31%.

Show work:

Final Answer: