1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

Number of Zn atoms = number mole × Avogadros number

Number of Zn atoms =   × 6.02×10^{23}

Answer = $\times 10^{24}$

2. Calculate the number of molecules in 11.5 mol of water (H_2O).

Number of H_2O molecules = number mole × Avogadros number

Number of H_2O molecules =   × $\times 10^{23}$

Answer = $\times 10^{24}$