

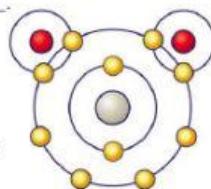
Chemical bonding

Name _____ Class _____ Date _____

1

A covalent bond usually forms between _____.

- A two metals
- B two metalloids
- C a metal and a nonmetal
- D two nonmetals



2

Using the **electron dot diagram** below, determine what kind of **bond** has formed between the carbon atom and the two oxygen atoms.

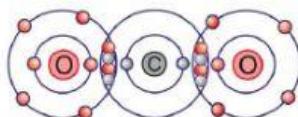
- A metallic
- B covalent
- C ionic
- D molecular



3

Using this diagram of carbon dioxide, determine how many **valence electrons** each **oxygen** atom has after bonding with the carbon atom.

- A 8
- B 6
- C 4
- D 2



4

In the diagram below, one **sodium** atom has bonded to one **chlorine** atom. What type of **bond** has formed between these two atoms?

- A metallic
- B covalent
- C ionic
- D molecular



5

This **metal** and **nonmetal** have formed an ionic bond. What is the name of the **compound** that has been formed?

- A sodium chlorine
- B sodium chloride
- C sodiumide chloride
- D sodiumide chlorine



6

When a group of **active metal atoms** are put together, a strong force of **attraction** called a(n) _____ develops between these atoms.

- A ionic bond
- B covalent bond
- C metallic bond
- D magnetic bond

7

The **diagram** below **shows** that aluminum has bonded to chlorine. The proper term for this **diagram** is a _____.

- A chemical formula
- B compound formula
- C mass formula
- D reaction formula



8

How many atoms make up the compound shown below?

- A 1
- B 2
- C 3
- D 4



9

The name of the compound shown in the **chemical formula** below is _____.

- A aluminum chlorine
- B chlorine aluminum
- C aluminum chloride
- D chloroaluminum



10

What is the **correct** formula for **two** molecules of sodium bromide?

- A NaBr_x2
- B 2NaBr
- C NaBr₂
- D 2xNaBr

