

## Mole Conversions: In Class Practice

1. How many grams is 2.0 mol of  $\text{H}_2\text{O}$ ?

What is the molar mass of  $\text{H}_2\text{O}$ ? \_\_\_\_\_ g/mol

$$\text{X} \quad \underline{\hspace{2cm}} \quad =$$

2. How many moles is 82.0 grams of  $\text{CH}_4$ ?

What is the molar mass of  $\text{CH}_4$ ? \_\_\_\_\_ g/mol

$$\text{X} \quad \underline{\hspace{2cm}} \quad =$$

3. How many moles is 56.04 grams of  $\text{N}_2$ ?

What is the molar mass of  $\text{N}_2$ ? \_\_\_\_\_ g/mol

$$\text{X} \quad \underline{\hspace{2cm}} \quad =$$

4. What is the mass of 1.50 moles of  $\text{CO}_2$ ?

What is the molar mass of  $\text{CO}_2$ ? \_\_\_\_\_ g/mol

$$\text{X} \quad \underline{\hspace{2cm}} \quad =$$

5. You have 17.0 g of  $\text{PH}_3$ . How many moles do you have?

What is the molar mass of  $\text{PH}_3$ ? \_\_\_\_\_ g/mol

$$\text{X} \quad \underline{\hspace{2cm}} \quad =$$