# EQUILIBRIUM EXPRESSIONS
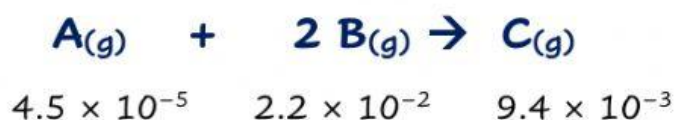
Q1 : Choose the correct answer :

For the reaction  $2A + B \rightarrow 2C$  the appropriate form for the equilibrium constant expression is :

a. $\dfrac{[A][B]^2}{[C]}$

b. $\dfrac{[A]^2 [B]}{[C]^2}$

C. $\dfrac{[C]^2}{[A]^2 [B]}$

d. $[A][B]^2[C]$

Q2 : Choose the correct answer :

For the following hypothetical equilibrium, what is the value of the equilibrium constant if the concentrations at equilibrium are as shown?

$$A_{(g)} \quad + \quad 2\,B_{(g)} \rightarrow C_{(g)}$$

$$4.5 \times 10^{-5} \quad 2.2 \times 10^{-2} \quad 9.4 \times 10^{-3}$$

a. 0.22            b. 9.9            c. $4.3 \times 10^5$            d. $2.3 \times 10^8$