How many neutrons are in each of the following isotopes:

Magnesium-24 has _____ neutrons

Mercury 201 has _____ neutrons

Osmium-190 has _____ neutrons

Radon-222 has _____ neutrons

Aluminum-27 has _____ neutrons

Platinum-195 has _____ neutrons

Name the isotope with the following combination of protons and neutrons.

58 protons and 82 neutrons = _____

53 protons and 74 neutrons = _____

20 protons and 20 neutrons = _____

82 protons and 125 neutrons = _____

Working with ISOTOPES

- Atomic Number = number of _____.
- Mass Number = number of _____+_____ in an atom of an isotope
- number of neutrons = _____ Number – _____ Number.

Fill in the table below with the appropriate information for each **isotope:**

Isotope	Atomic Number	Mass Number	# of protons	# of neutrons
Germanium-73				
		16	8	
Gold-197				
		9	4	
Hydrogen-3				
Plutonium-244				
Potassium-39				
Bromine-80				

Remember: Atoms of the **SAME ELEMENT** with **DIFFERENT NUMBERS OF NEUTRONS** are called _____.