____ 44. In chemical reactions, what does the principle of conservation of mass mean?
a. Matter is not changed.
b. The total mass of the reactants is less than the total mass of the products.
c. The total mass of the reactants is greater than the total mass of the products.
d. Matter is not created or destroyed.

____ 45. Water vapor in the air turns to liquid water in the form of rain. This is an example of a
a. chemical change.
b. chemical equation.
c. physical change.
d. chemical formula.

____ 46. A material used to increase the rate of a chemical reaction is a(n)
a. enzyme.
b. fuel.
c. catalyst.
d. inhibitor.

____ 47. A reaction that absorbs energy in the form of heat is described as
a. combustion.
b. unbalanced.
c. exothermic.
d. endothermic.

____ 48. The substances listed on the left side of a chemical equation are the
a. coefficients.
b. reactants.
c. products.
d. precipitates.

____ 49. In an equation, numbers often appear in front of a chemical formula. These numbers tell you the
a. number of molecules in each atom in the reaction.
b. number of molecules or atoms of each substance in the reaction.
c. number of atoms in each molecule in the reaction.
d. number of elements in the reaction.

____ 50. Every chemical reaction involves a change in
a. state.
b. energy.
c. concentration.
d. mass.