



Avagadro's Number = 6.02×10^{23} atoms/mol

1 mol of a gas at STP occupies 22.4 L

The answer of question is given at the bottom of the table.

Drag the correct answer and **Drop** in the box against the question.

| S.No. | QUESTION | Answer |
|-------|---|--------|
| 1. | How many atoms of Oxygen are there in 18g of | |
| 2. | How many atoms of Hydrogen are there in 18g of water? | |
| 3. | How many molecules of H_2O are there in 18g of | |
| 4. | What is the mass of 1 mole of O_2 ? | |
| 5. | What is the mass of 1 molecule of O_2 ? | |
| 6. | What is the mass of 2 mol of H_2SO_4 ? | |
| 7. | What is the density of O_2 at STP? | |
| 8. | 3 L of a gas weighs 2 g. What is its molecular mass? | |
| 9. | What volume does 22g of CO_2 at STP occupy? | |
| 10. | How many atoms of Hydrogen are in 67.2 L of H_2 at STP? | |

6.02×10^{23} 196g 1.204×10^{24} 1.43g/L 6.02×10^{23} 3.612×10^{24}
 14.9g/mol 32g 11.2L $5.32 \times 10^{-23}g$