Study Guide Chapter Four - Bonding

magnesium oxide	2	a negative ion	2
oppositely charged ions	covalent	covalent bond	polyatomic ions
positive ion	sodium sulfide	when dissolved in water	neutral
an atom or group of atoms that has become electrically charged	they have low melting points	electrons are transferred	

1. What is an ion? _____

_____

2. When an atom loses an electron, it becomes a _____
3. When an electron is transferred from a sodium atom to a chlorine atom, the chlorine atom becomes _____
4. An ionic bond is the attraction between _____
5. An ionic bond is formed when _____
6. Ionic compounds are electrically _____
7. How many chlorine ions are needed to cancel the 2+ charge of magnesium in magnesium chloride ($MgCl_2$). _____
8. Ions that are made of more then one atom are examples of

_____

9. What is the chemical name for the compound with the formula Na_2S_____
10. The ionic compound MgO is called _____
11. This is NOT a characteristic property of ionic compounds

_____

12. In what form can an ionic compound conduct electricity?

_____

13. A chemical bond that forms when two atoms share electrons is called a(n)

_____

14. Bonds that form between two nonmetals are usually _____
15. How many covalent bonds can oxygen form? _____

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