

Writing Formulas and Naming Compounds

Directions: Answer the following questions in the spaces provided. Refer to the periodic table for help.

1. Define an oxidation number. Is the number of electrons an atom will **lose** **gain** to become a positive ion or **lose** **gain** to become a negative ion and have a full outer energy level of _____ electrons
2. What is the usual oxidation number of oxygen? Of hydrogen? O H
3. What is the sum of all the oxidation numbers in any compound? use a number

Directions: Use the periodic table to identify the oxidation numbers of the elements in each group.

Group	1	2	6	7	8
Oxidation number					

Directions: Write the formulas for the following compounds. Use the periodic table for help.

Symbols with oxidation numbers

Formula

10. copper(II) sulfate _____
11. calcium chloride _____
12. iron(II) oxide _____
13. copper(I) oxide _____
14. sodium sulfide _____
- 15 magnesium bromide _____

Directions: complete the following table, providing the name of the compound and the total number of atoms in each formula given.

Formula	Name	Number of Atoms
16. (NH ₄) OH		
17. (NH ₄) Cl		
18. Ag ₂ O		
19. K ₂ (SO ₄)		
20. Ca (NO ₃) ₂		
21. Na ₂ S		