

NAME: _____

CLASS&SEC

Chemical reactions one-mark test

1. A reaction in which a single product is formed from two (or) more reactants is known as **reaction.**

a) Displacement b) oxidation c) combination d) decomposition

2. The substance which is formed after completion of the reaction is

3. Destruction of ammonium dichromate is an example of _____ reaction.

a) Decomposition b) combination c) displacement d) reduction

4. The reaction in which a more reactive element displaces a less reactive element from its compound is called

a) Oxidation b) reduction c) combination d) displacement

5. During the chemical reaction the insoluble substance is formed. It is technically termed as

a) Product b) precipitate c) slag d) matte

6. The reaction in which exchange of ions between two reactants occur, leading to the formation of two different products is called _____ reaction.

- a) Double Displacement**
- b) double decomposition**
- c) both A and B**
- d) displacement**

7. The following reaction is an example of a
 $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

(i) displacement reaction (iii) redox reaction	(ii) combination reaction (iv) neutralisation reaction
(a) (i) and (iv)	(b) (ii) and (iii)
(c) (i) and (iii)	(d) (iii) and (iv)

8. Which is less reactive?

9. Which among the following is(are) double displacement reaction(s)?

(i) $\text{Pb} + \text{CuCl}_2 \rightarrow \text{PbCl}_2 + \text{Cu}$ (iii) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$	(ii) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$ (iv) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
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(a) (i) and (iv) (b) (ii) only
(c) (i) and (ii) (d) (iii) and (iv)

10. When lead powder is added to copper chloride solution, a displacement reaction occurs and solid copper is formed. Why does the displacement reaction occur?

(lead is more reactive than copper / copper is more reactive than lead)

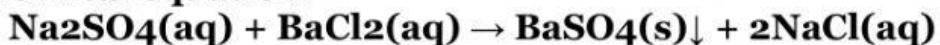
11. $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl}(\text{g})$ is a

a. Decomposition Reaction **b. Combination Reaction**
c. Single Displacement Reaction **d. Double Displacement Reaction**

12. Photolysis is a decomposition reaction caused by

a. heat b. electricity c. light d. mechanical energy

13. The chemical equation



represents which of the following types of reaction?

- a. Neutralisation**
- b. Combustion**
- c. Precipitation**
- d. Single displacement**

14. Which of the following represents a precipitation reaction?

a. $A(s) + B(s) \rightarrow C(s) + D(s)$ b. $A(s) + B(aq) \rightarrow C(aq) + D(l)$
c. $A(aq) + B(aq) \rightarrow C(s) + D(aq)$ d. $A(aq) + B(s) \rightarrow C(aq) + D(l)$

15. Which of the following is not an “element + element → compound” type reaction?

a. $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$ b. $2\text{K(s)} + \text{Br}_2\text{(l)} \rightarrow 2\text{KBr(s)}$
c. $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$ d. $4\text{Fe(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Fe}_2\text{O}_3\text{(s)}$