Activity 1 : Writing the Thermochemical equation:

1. Write the thermochemical equation of Combustion of sucrose ($C_{12}H_{22}O_{11}$) forming carbon dioxide and liquid water , the energy released is 5644 KJ/ mol.

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2. Write the thermochemical equation of combustion of methane (CH_4)

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3. Write the thermochemical equation of combustion of methane (C_2H_5OH) ,

   $\Delta H = -1367$ KJ

Activity 2 : Identify which of the reaction is endothermic reaction and which one is exothermic reaction :

EQUATION	TYPE OF REACTION	EQUATION	TYPE OF REACTION
$C_3H_{8(g)} \rightarrow C_3H_{8(l)}$		$C_{10}H_{8(S)} \rightarrow C_{10}H_{8(l)}$	
$CO_{2(s)} \rightarrow CO_{2(g)}$		$H_2O_{(l)} \rightarrow H_2O_{(S)}$	