

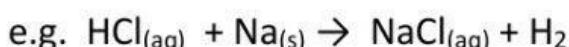
REACTIONS OF ACIDS AND BASES

NEUTRALIZATION

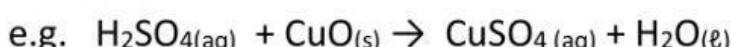
When aqueous solutions of acids react with bases, neutralization takes place and salts are formed.

The four neutralization reactions that form salts are as follows:

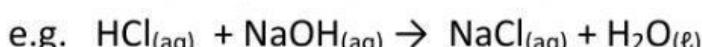
1. acid + metal \rightarrow salt + hydrogen



2. acid + metal oxide \rightarrow salt + water



3. acid + metal hydroxide \rightarrow salt + water



4. acid + metal carbonate \rightarrow salt + water + carbon dioxide



The type of salt that forms depends on the acid:

- hydrochloric acid gives a chloride
- nitric acid gives a nitrate
- sulphuric acid gives a sulphate
-

You'll need to make sure you know your compound ions list for this section!

Complete the following table by filling in the **formula, charge or name** of the compound ion.

sulphate	SO_4
carbonate	
Nitrate	
Phosphate	
	OH^{-1}
Ethanoate/acetate	
ammonium	

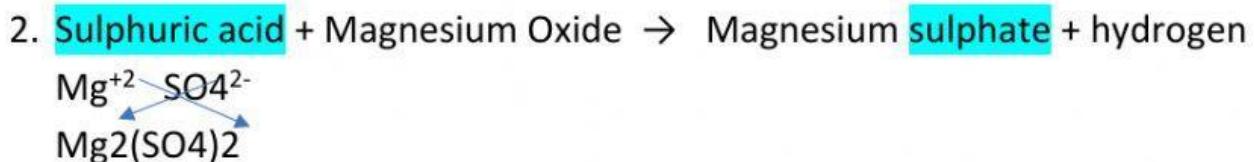
Let's practice forming some salts:



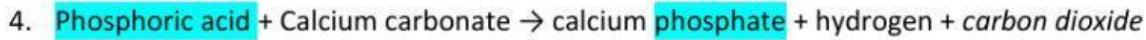
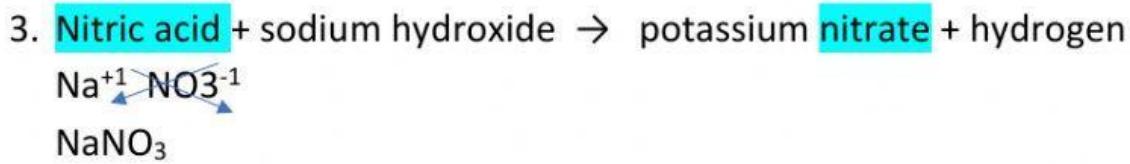
The salt that forms, comes from the metal and the negative part of the acid.
 K^+ and Cl^-

Don't forget to cross multiply the charges to get your salt's formula
 $\text{K}^{+1} \text{Cl}^{-1}$ *remember the signs (+ or -) don't go with the number when you cross multiply.

K_1Cl_1 - but we never write the 1's
∴ KCl



But we then need to simplify this formula and 'divide' by 2
 MgSO_4



Remember that a metal carbonate and an acid form an additional CO_2 that the others don't form



$\text{Ca}_3(\text{PO}_4)_2$ Remember to write the compound ion in brackets if it has been multiplied by a number

*There are 2 acids where the salt has to be written in a different way.

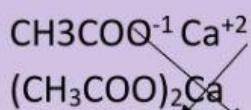
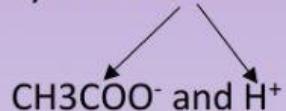
They are: acetic acid and oxalic acid

CH_3COOH and $(\text{COOH})_2$

Normally when you write the salt, the metal is written first and then the negative part of the acid. However with these 2 salts, the acid part will be written first and then the metal

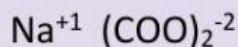
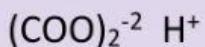
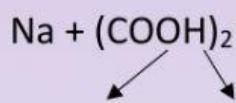
Eg. $\text{CH}_3\text{COOH} + \text{Ca} \rightarrow$

Firstly the CH_3COOH splits up into



Thus $(\text{CH}_3\text{COO})_2\text{Ca}$

The same happens to oxalic acid:



Thus $(\text{COO})_2\text{Na}_2$

The gases that are produced in the reactions above can be identified by doing the following tests:

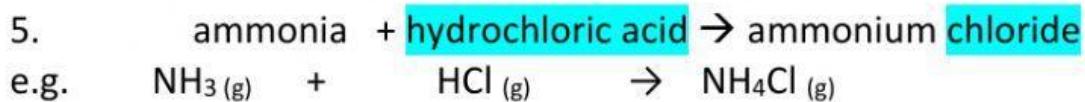
hydrogen gas : a glowing splint will ignite the gas with a 'pop' or a 'squeal' in the presence of the gas.

carbon dioxide : turns clear lime water milky (it also causes a glowing splint to extinguish)

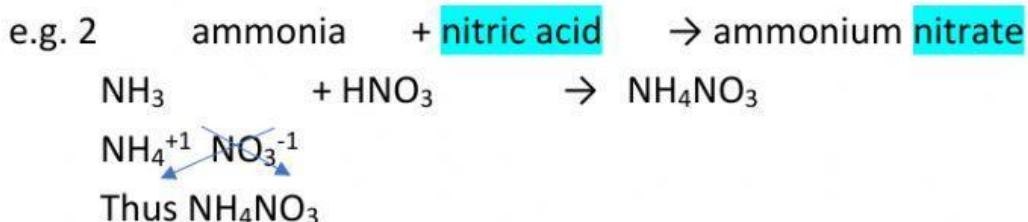
Note: not all metals react with acid, only those that are more reactive than hydrogen. These include K, Na, Li, Ca, Mg, Al, Zn and Fe.

REACTION OF AMMONIA WITH HYDROCHLORIC ACID

This is also a neutralization reaction, but is also a synthesis reaction.



In this case the salt that forms is always an ammonium salt



QUESTION 1

Complete the following word equations:

- 1.1 acid + carbonate → salt + _____ + _____
- 1.2 acid + metal → salt + _____
- 1.3 acid + metal oxide → salt + _____
- 1.4 acid + base → salt + _____

QUESTION 2

Name the salt that forms when: (two words)

- 2.1 HCl reacts with KOH
- 2.2 sulphuric acid reacts with calcium metal
- 2.3 HNO₃ reacts with CaCO₃
- 2.4 ethanoic acid react with sodium oxide

QUESTION 3 *write subscripts as normal number and write balancing number in their own blocks {if no balancing no is required, ensure to write a 1}

Complete the following reactions and balance the equations:

- 3.1 HCl + Mg → +
- 3.2 HNO₃ + Zn → +
- 3.3 H₂SO₄ + Ca → +
- 3.4 HCl + MgO → +
- 3.5 HNO₃ + Li₂O → +
- 3.6 H₂SO₄ + Na₂O → +
- 3.7 HCl + Ca(OH)₂ → +
- 3.8 H₂SO₄ + Mg(OH)₂ → +
- 3.9 HNO₃ + CaCO₃ → + +
- 3.10 CH₃COOH + MgO → +
- 3.11 H₂SO₄ + Zn → +
- 3.12 HCl + NH₃ →

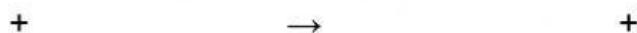
QUESTION 4

Write balanced chemical reactions for the following reactions:

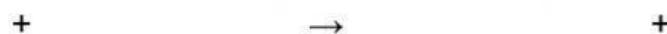
***subscripts can be written as normal numbers**

***balancing numbers must be put into separate boxes in front of the substance**

4.1 zinc metal reacts with sulfuric acid.



4.2 potassium hydroxide is added to hydrochloric acid



4.3 sodium oxide reacts with nitric acid.



4.4 ammonia gas reacts with hydrochloric acid.



4.5 calcium reacts with phosphoric acid.

