Name:
Stoichiometry

1. ___4___ **Fe** + ___3___ $\mathbf{O_2}$ → ___2___ $\mathbf{Fe_2O_3}$

a) How many **moles** of Fe would be needed to react with 3.82 **moles** of O_2?

_____ = mol Fe

b) What **mass** of Fe_2O_3 can be produced from 13.5 **grams** Fe?

_____ = g Fe_2O_3

c) How many **grams** of O_2 are needed to produce 34.7 **g** of Fe_2O_3?

_____ = g O_2

d) What **moles** of Fe_2O_3 can be produced from 1.35 **mol** Fe?

_____ = mol Fe_2O_3

2. ____$KClO_3$ ____KCl + ____O_2

When 62.0 g of Potassium chlorate decomposes, how many grams of KCl will be formed?

_____ =

LIVEWORKSHEETS