

Persamaan Pengionan / Ionisation Equation

Ionisation of acid <i>Pengionan asid</i>			Ionisation of alkali <i>Pengionan alkali</i>		
HNO_3 Nitric acid <i>Asid nitrik</i>	$\longrightarrow$	 Hydrogen ion <i>Ion hidrogen</i>	+	 Nitrate ion <i>Ion nitrat</i>	
H_2SO_4 Sulphuric acid <i>Asid sulfurik</i>	$\longrightarrow$	 Hydrogen ion <i>Ion hidrogen</i>	+	 Sulphate ion <i>Ion sulfat</i>	
H_3PO_4 Phosphoric acid <i>Asid fosforik</i>	$\longrightarrow$	$3\text{H}^+(\text{aq/ak})$ Hydrogen ion <i>Ion hidrogen</i>	+	 Phosphate ion <i>Ion fosfat</i>	
CH_3COOH Ethanoic acid <i>Asid etanoik</i>	$\rightleftharpoons$	 Ethanoate ion <i>Ion etanoat</i>	+	 Hydrogen ion <i>Ion hidrogen</i>	
Ammonia <i>Ammonia</i>	$\xrightleftharpoons{\text{H}_2\text{O}}$	 Ammonium ion <i>Ion ammonia</i>	+	 Hydroxide ion <i>Ion hidroksida</i>	
Sodium hydroxide <i>Natrium hidroksida</i>	$\text{NaOH} \longrightarrow$	 Sodium ion <i>Ion natrium</i>	+	 Hydroxide ion <i>Ion hidroksida</i>	
Potassium hydroxide <i>Kalium hidroksida</i>	$\text{KOH} \longrightarrow$	 Potassium ion <i>Ion kalium</i>	+	 Hydroxide ion <i>Ion hidroksida</i>	
Barium hydroxide <i>Barium hidroksida</i>	$\text{Ba(OH)}_2 \longrightarrow$	 Barium ion <i>Ion barium</i>	+	 Hydroxide ion <i>Ion hidroksida</i>	