Learning Target: I can calculate percent composition by mass.

Calculating Percent Composition by Mass Video Notes

1. **What is the percent composition by mass of Hydrogen in H_2O?**

$$\text{percent composition by mass:} \text{(percent by mass)} \quad \frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$$

Calculate molar mass of H_2O -   H = 2 x (_____) + H = 1 x (____) = molar mass of H_2O = _____

Write down the notes at the 3-minute mark of the video to solve the problem.

What is the percent composition by mass of Hydrogen in H_2O? = _____

2. **Calculate the percent by mass of Iron in $FeCl_3$**

Calculate the molar mass of **$FeCl_3$** = _____   Calculate the molar mass of Fe = _____

Input your numbers in the following formula to solve. $\frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$

The percent by mass of Iron in $FeCl_3$ = _____

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LIVEWORKSHEETS

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3. Determine the percent by mass of Carbon in $C_6H_{10}O_4$.

Calculate the molar mass of $C_6H_{10}O_4$ = _____   Calculate the molar mass of Carbon = _____

Input your numbers in the following formula to solve.

$$\frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$$

The percent by mass of Carbon in $C_6H_{10}O_4$ = _____

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