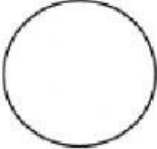
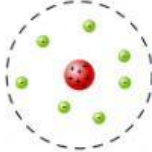
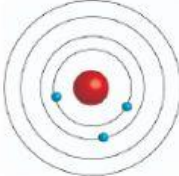
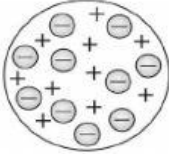


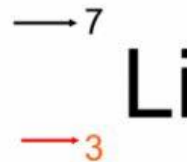
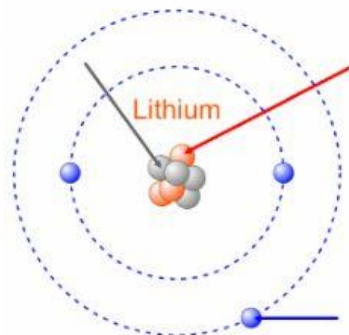
Atomic Structure and History Basics

Match the atomic model or discovery proposed by each scientist by placing the appropriate letter in the box. **A) Rutherford** **B) Bohr** **C) Thomson** **D) Dalton**

Planetary model	
Plum pudding model	
Gold foil Experiment	
First Atomic Theory	

Label the following parts of the atom with the terms:

proton neutron electron mass number atomic number



Complete the following Tables

Subatomic Particle	Charge (+, - or 0)	Mass (amu) (1 or 0)	Location (inside or outside)
proton			
neutron			
electron			

Questions

- The number of protons is equal to the _____.
a. Mass number b. Atomic number c. the number of neutrons
- An element has the atomic number of 14. What is the element?
a. Nitrogen b. Silicon b. Phosphorus
- Isotopes have a different number of _____.
a. Protons b. Electrons c. neutrons
- Which of these atoms has the greatest number of neutrons in its nucleus?
a) $^{56}_{25}\text{Mn}$ c) $^{57}_{27}\text{Co}$
b) $^{56}_{26}\text{Fe}$ d) $^{56}_{28}\text{Ni}$

Complete the following table:

Element	Symbol	Atomic Number	No. of protons	No. of Neutrons	Mass Number	No. of electrons
	$^{35}_{17}\text{___}$	17			35	17
	$^{\text{---}}_1\text{H}$	1		0		
sodium	^{23}Na		11			
	$^{31}_{15}\text{P}$			16		