

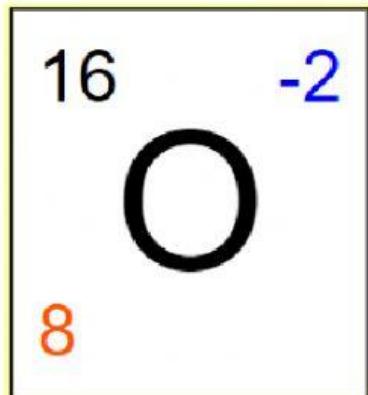
QUIZIZZUnit 1: Atomic Structure
30 Questions

NAME : _____

CLASS : _____

DATE : _____

1.



What is Oxygen's atomic number?

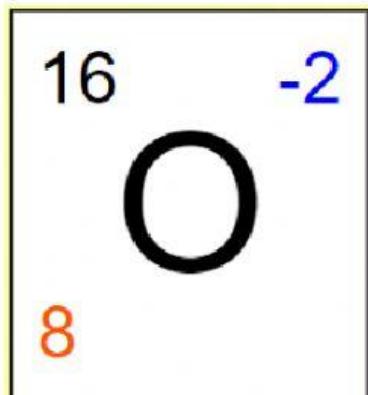
-2

8

10

16

2.



What is this atom's atomic mass?

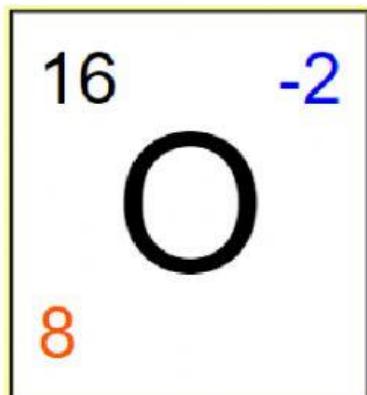
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3.



What is this atom's charge?

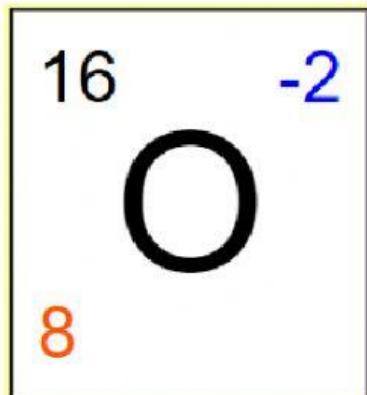
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4.



How many neutrons does this atom have?

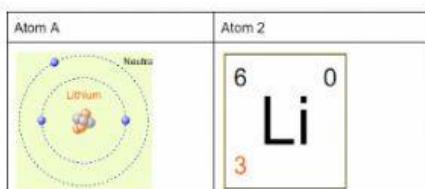
10

8

-2

16

5.



Determine the relationship between the following atoms:

The same atom

A different element

Isotopes of the same element

6. What is the charge of electrons?

Electrons have no charge (0)

Electrons are positively (+) charged

Electrons are negatively (-) charged

7. What is the mass of a proton?

0

1

-1

8. Who discovered electrons?

Alexander the Great

JJ Thomson

Niels Bohr

9. What is the charge of a proton?

-

0

+

10. What is the charge of a neutron?

-

+

0

11. What is the mass of a neutron?

0

-1

1

12.



How many NEUTRON are there in this Hydrogen?

0 2

1

13. What makes up the nucleus?

Electrons+ Protons

Protons + Neutrons

Neutrons + Electrons

14. What are three subparticles of an atom?

Proton, Neutron, Electron

One Oxygen and two Hydrogen. H_2O

Hydrogen, Oxygen, Helium

15. Where do electrons found?

Electrons are not moving

Around the nucleus

In the center of the atom

16. Where are protons found?

Around the nucleus

In the center of the atom

17. Subatomic particles with a negative charge

Electrons	Quarks
Protons	Neutrons

18. Subatomic particles with a positive charge

neutrons	atomic mass
isotopes	protons

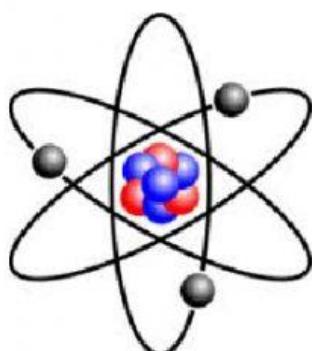
19. Subatomic particles that are neutral in charge

Neutrons	Protons
Electrons	Nucleus

20. A tiny but very dense, positively charged portion of the atom that holds most of the atomic mass

Electron Cloud	Orbitals
Nucleus	Electron Shells

21.



What is one way Neutrons are different from electrons?

Both are in the nucleus(center of atom)

Electrons are inside nucleus and neutrons are outside the nucleus.

neutrons have no charge, while electrons have a negative charge

they both are negative

28. What is the mass of a carbon atom that has 6 protons, 7 neutrons, and 6 electrons?

13

7

6

19

29. Which subatomic particle determines the identity of an element?

neutrons

protons

electrons

30. An atom has:

8 Protons

7 Neutrons

6 Electrons

What is the mass number?
