

Atomic Structure Worksheet (with ions)

Complete entire table (including the first column) **without using a periodic table:**

$p + n = \text{mass number}$

$p = \text{atomic number}$

$$\begin{matrix} A \\ Z \end{matrix} X$$

$p - e = \text{charge}$

When in this form ... assume it is neutral

Name - mass number

Element	Atomic #	Mass #	${}_1^1p^+$	${}_0^1n^0$	${}_{-1}^0e^-$
${}_{13}^{27}\text{Al}^{3+}$		27			
Lead - 209	82				
${}_6^{12}\text{C}^{4+}$			6	7	
${}_{56}^{139}\text{Ba}^{2+}$	56				
${}_{36}^{84}\text{Krypton}$					36
${}_{17}^{35}\text{Cl}^{1-}$		35			
${}_{11}^{23}\text{Na}^{1+}$			11	12	
Nitrogen - 15	7				
Selenium - 80	34				
Strontium - _____	38	88			
Tin - 118	50				
Rubidium - 86	37				
Xenon - 131					54
Fluorine - _____		19	9		
${}_{19}^{39}\text{K}^{1+}$					