INTRODUCTION TO REDOX CHEMISTRY

Complete the sentences below.

1) Oxidation occurs when a substance _____ electrons and _____ in oxidation state.

2) Reduction occurs when a substance _____electrons and _____in oxidation state.

Determine and then write the oxidation state of the underlined element in the space provided. You must include the "+" or "-" sign in front of the oxidation number.

Example:

$\underline{Cr}_2O_7^{2-}$ Oxidation State: +6

3) $\underline{S}O_2$ Oxidation State: _____

4) $H\underline{N}O_3$ Oxidation State: _____

5) $\underline{Cr}O_4^{2-}$ Oxidation State: _____

6) $Na\underline{Cl}O_3$ Oxidation State: _____

7) $\underline{Mn}O_2$ Oxidation State: _____

8) $\underline{V}O_4^{3-}$ Oxidation State: _____

9) $\underline{Mn}O_4^{2-}$ Oxidation State: _____

10) $\underline{Cr}_2O_3$ Oxidation State: _____