

Learning Target: I can describe and explain the physical and chemical properties of intermolecular forces.



Bozeman Science Intermolecular Forces Video Notes

1. Intermolecular forces are forces between _____.
2. What are the intermolecular force properties of solids? _____
3. What are the intermolecular force properties of gases? _____
4. What are the intermolecular force properties in biological systems? _____
5. What is the difference between intermolecular and intramolecular forces? _____
6.

Draw an example of strong intermolecular forces.	Draw an example of weak intermolecular forces.

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7. What properties are affected by intermolecular forces? _____

8. Boiling point is how much _____ energy is needed to change a _____ into a _____. The greater the intermolecular forces, then the _____ it will be to go from a liquid to a gas.

9. Surface tension is like a _____ over a liquid. Out of water, alcohol, and mercury, which has the greatest surface tension? _____ Why? _____

10. Capillary action is movement of a _____ up a tube. The thinner we make the tube then _____

What is an everyday example of capillary action? _____

11. Vapor pressure is measured in a _____ system. Vapor pressure is the amount of liquid that goes from a liquid phase to a _____ phase. The higher the vapor pressure, the higher the _____ because there are _____ intermolecular forces. THINK ABOUT WHEN YOU SHAKE UP A SODA CAN!!!

12. Miscibility is how easy it is for two materials to _____ together. Why does water not mix well with gasoline? _____

Why does water mix well with hydrochloric acid? _____

13. Why is it easy for NaCl (Salt) to mix well with water? _____

14. What is the formula for ideal gas? _____ What is the difference between ideal gas and real gas? _____

15. PV / RT for a real gas is _____. What happens when you change the temperature for an ideal gas? _____

16. What makes an enzyme work? (Explain using intermolecular forces) _____