Unit 4 Study Guide General

Molecular compounds that dissolve in water do not conduct electricity because no are present.

When an atom loses an electron, it becomes a

- ○ a neutral atom.
- ○ b neutral ion.
- ○ c positive ion.
- ○ d negative ion.

When an atom gains an electron is become a

- ○ a negative ion
- ○ b neutral atom
- ○ c neutral ion
- ○ d positive ion

Select all of the properties of Ionic Compounds

- ☐ a Do not break into ions when dissolved
- ☐ b Lack orderly crystal patterns
- ☐ c Low melting points
- ☐ d Usually liquids or gases at room temperature
- ☐ e Conduct electricity when dissolved
- ☐ f Are called molecules
- ☐ g Do not conduct electricity when dissolved
- ☐ h High melting points
- ☐ i Break into ions when dissolved
- ☐ j Usually solids at room temperature
- ☐ k Orderly crystal patterns

The sum of the charges on an ionic compound is always _____

An Ionic compound is the result of a _____ bonding with a _____.

- ○ a metal, metal
- ○ b non-metal, non-metal
- ○ c noble gas, metal
- ○ d metal, non-metal

The ions in an ionic compound form regular patterns because

- ○ a every ion is attracted to ions near it with an opposite charge.
- ○ b every ion shares an electron with the ion next to it.
- ○ c ions are always found in a 1:1 ratio.
- ○ d each ion has a square shape.

What is an ion?

- ○ a an atom that doesn't have protons
- ○ b an atom or group of atoms that has become electrically charged
- ○ c the opposite of an electron
- ○ d an atom that doesn't have neutrons

LIVEWORKSHEETS