

Unit 4 Study Guide General

Molecular compounds that dissolve in water do not conduct electricity because no _____ are present.

When an atom loses an electron, it becomes a _____

- a neutral atom.
- c positive ion.
- b neutral ion.
- d negative ion.

When an atom gains an electron is become a _____

- a negative ion
- b neutral atom
- c neutral ion
- d positive ion

Select all of the properties of Ionic Compounds

- a Do not break into ions when dissolved
- b Lack orderly crystal patterns
- c Low melting points
- d Usually liquids or gases at room temperature
- e Conduct electricity when dissolved
- f Are called molecules
- g Do not conduct electricity when dissolved
- h High melting points
- i Break into ions when dissolved
- j Usually solids at room temperature
- k Orderly crystal patterns

The sum of the charges on an ionic compound is always _____

An Ionic compound is the result of a _____ bonding with a _____.

- a metal, metal
- b non-metal, non-metal
- c noble gas, metal
- d metal, non-metal

The ions in an ionic compound form regular patterns because

- a every ion is attracted to ions near it with an opposite charge.
- b every ion shares an electron with the ion next to it.
- c ions are always found in a 1:1 ratio.
- d each ion has a square shape.

What is an ion?

- a an atom that doesn't have protons
- b an atom or group of atoms that has become electrically charged
- c the opposite of an electron
- d an atom that doesn't have neutrons