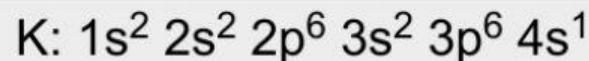


PERIODICITY: SIZES

ANSWER ALL QUESTIONS.



Which group does they belong to?

same group

isoelectronic

same period

Which atom is larger?

Na

K



Which group does they belong to?

same group

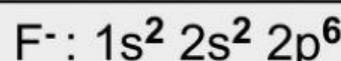
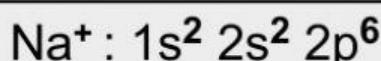
isoelectronic

same period

Which atom is larger?

Na

Mg



Which group does they belong to?

same group

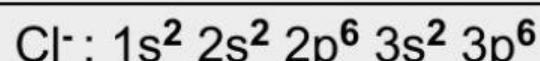
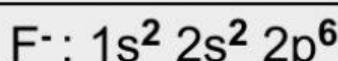
isoelectronic

same period

Which atom is larger?

Na⁺

F⁻



Which group does they belong to?

same group

isoelectronic

same period

Which atom is larger?

F⁻

Cl⁻

PERIODICITY: SIZES

ANSWER ALL QUESTIONS.



What happen when two electron are added to O atom?

electron – electron repulsion decrease

Size of O atom is smaller than size of O^{2-} ion.

electron – electron repulsion increase

Size of O atom is larger than size of O^{2-} ion.



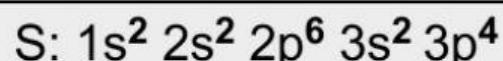
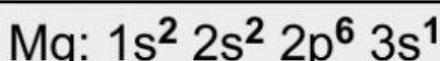
What happen when an electron is removed from K atom?

electron – electron repulsion decrease

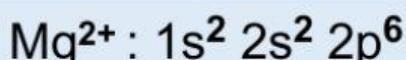
Size of K atom is smaller than size of K^+ ion.

electron – electron repulsion increase

Size of K atom is larger than size of K^+ ion.



Both atom are in period 3. Compare the size of S^{2-} to Mg^{2+}



Mg^{2+} has more shell than S^{2-}

Mg^{2+} has more shielding effect than S^{2-}

Nucleus attraction towards valence electrons in Mg^{2+} is stronger than S^{2-}

Mg^{2+} is smaller than S^{2-}



S^{2-} has more shell than Mg^{2+}

S^{2-} has more shielding effect than Mg^{2+}

Nucleus attraction towards valence electrons in S^{2-} is stronger than Mg^{2+}

Mg^{2+} is larger than S^{2-}