



Roatan Bilingual School
10th Grade Chemistry Evaluation
SEMESTER RECUPERATION

Name: _____

Date: _____

Ms. April Brown

100%

Read each question or exercise carefully. RESPOND IN YOUR OWN WORDS. to show your work and leave all evidence possible.

TRUE OR FALSE

Read each statement carefully. Write a T on the line next to it if it is true. Write an F if it is false and EXPLAIN WHY.

1. ____ The formula for Carbon Monoxide is CO₂.
2. ____ KBr is Potassium Baromdiide.
3. ____ NaF is a covalent compound.
4. ____ ZnI is an ionic compound.
5. ____ Pressure is a force or characteristic that can be measured in gasses.

The unit used to measure pressure is calories.

6. ____ H₂O is a molecule.
7. ____ Ne is a noble gas.
8. ____ H₂O and N₂ are examples of matter that are in plasma state.
9. ____ Temperature is measured ONLY in Celsius.
10. ____ Mass is usually measured in L.
11. ____ CaCO represents Calcium Nitrate
12. ____ Calories are used to measure temperature.

MULTIPLE CHOICE

Circle the best answer. If your choice is "none of the above" explain why.

1. This is a polyatomic compound

- a. Acetate
- b. NaNO_3
- c. K_3PO_4
- d. All of the above

2. Name KOH

- a. Cerium Oxide
- b. Kallium Oxihydride
- c. Potassium Hydroxide
- d. None of the above _____

3. What does the subscript in $\text{Al}_2(\text{SO}_4)_3$ represent?

- a. There is no subscript
- b. This is a formula for a different compound
- c. This is a covalent compound
- d. None of the above _____

4. Equation for a chemical reaction in which the number of atoms for each element in the reaction and the total charge is the same for both the reactants and the products

- a. Potassium, Copper, Carbon
- b. Calcium, Magnesium, Silver
- c. synthesis reaction
- d. None of the above _____

5. Avogadro's number

- a. 1×10^23
- b. 6.02×10^23
- c. 62.0×10^23
- d. None of the above _____

6. Unit used to measure the amount or quantity of substances in a chemical reaction.

- a. Liters
- b. Moles
- c. Atoms
- d. None of the above _____

7. (2.016 g/mol) is the molar mass of:

- a. N_2
- b. H_2

- c. Cl_2
- d. None of the above _____

8. A state of matter in which particles are NOT tightly packed together.

- a. liquid
- b. plasma
- c. solid
- D. None of the above _____

9. In the equation $\text{CO}_2 + \text{H}_2\text{O} + \text{ATP} \rightarrow \text{O}_2 + \text{C}_6\text{H}_{12}\text{O}_2$, which compound is part of the product?

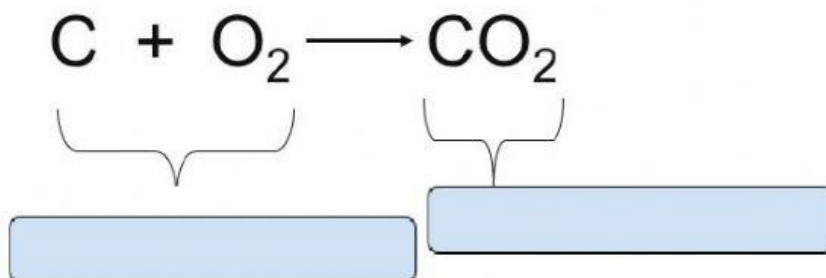
- a. H_2O
- b. CO_2
- c. ATP
- D. None of the above _____

9. In the equation $\text{CO}_2 + \text{H}_2\text{O} + \text{ATP} \rightarrow \text{O}_2 + \text{C}_6\text{H}_{12}\text{O}_6$, which compound is part of the reactants?

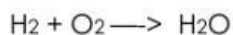
- a. H_2O
- b. $\text{C}_6\text{H}_{12}\text{O}_6$
- c. O_2
- D. None of the above _____

PRACTICAL EXERCISES

1. Name the parts of an equation



2. Balance the equations



3. Write a synthesis reaction: N, H, 2NH₃ (balance the equation)

4. Write a decomposition reaction: Zn, S, ZnS

5. Name the compounds



6. Write the formula

Glucose=

Lead Nitrate

Calcium Hydroxide

7. Tell whether single displacement or double displacement



8. Point out the subscripts.



Reactants

Iron

Hydrogen

Oxygen

Products

Iron

Oxygen

Hydrogen

9. Point out the coefficients

**Reactants**

Iron

Hydrogen

Oxygen

Products

Iron

Oxygen

Hydrogen

1. Describe the 4 states of matter discussed in class. 1pt

2. Name the following compounds. 4pts

Ca(OH)

HCL

NaCl

3. Convert the temperatures. Use the conversion factors. 2pts

110F to C

50C to F

4. Convert the amounts. Use the conversion factors. 3pts

1 L = 1000mL 1 kg = 1000g 1hr = 60min

360min to hr

600 mL to L

8900g to Kg

5. Balance the equations. 2pts



6. Find the molar mass. 4pts

Ethane C₂H₆

KCl

7. Covert from moles to grams. 2pts

7.5 moles of CO₂

10 moles of H₂

8. Convert from grams to moles. 2pts

454 g of NH₃

70g of NO₂

9. Identify the formula units. 2pts

0.533 mol NH₄

1. 65 g of H₂

CELSIUS TO FAHRENHEIT

$$T_F = \left(\frac{9}{5} T_C \right) + 32$$

FAHRENHEIT TO CELSIUS

$$T_C = \frac{5}{9} (T_F - 32)$$